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the study of quantitative relationships between the amounts of reactants used and products formed in a chemical reaction. mole ratio. ratio between numbers of moles of any two substances in a balanced chemical equation. limiting reactant.

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STUDY GUIDE FOR CONTENT MASTERY Section 12.2 Stoichiometric Calculations In your textbook, read about mole-to-mole conversion. Read the following passage and then solve the problems. In the equation that follows each problem, write in the space provided the mole ratio that can be used to solve the problem.

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12.2 Chemical Calculations. • In chemical calculations, mole ratios are used to convert between moles of reactant and moles of product, between moles of reactants, or between moles of products. • In a typical stoichiometric problem, the given quantity is first converted to moles.

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The reactant that is not completely used up in a reaction. The maximum amount of product that could be formed from given amounts of reactants. The amount of product that actually forms when the reaction is carried out in the laboratory. The ratio of the actual yield to the theoretical yield expressed as a percent.

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The study of the quantitative relationships between the amounts of reactants used and the amounts of products formed by a chemical reaction is called stoichiometry. Stoichiometry is based on the law of conservation of mass. In any chemical reaction, the mass of the productsA9han the mass of the reactants.

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____ 1. The study of the quantitative relationships between the amounts of reactants used and the amounts of products formed by a chemical reaction is called stoichiometry. ____ 2. Stoichiometry is based on the law of conservation of mass. ____ 3.

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Stoichiometry-the quantitative relationship between the reactants and products in a chemical reaction. Example: 2 Ca 3 (PO 4) 2 + 6SiO 2 + 10C 1P 4 + 6CaSiO 3 + 10CO From the above equation, we can determine that for every two moles of Ca 3 (PO 4) 2 that react, one mole of P 4 will be produced. Therefore, the molar ratio between Ca 3 (PO 4) 2 and P 4 is 2:1.

CHEMISTRY I (TESC 141) STUDY GUIDE - UW Tacoma

Name Date CHAPTER Section 11.1 continued In your textbook, read about mole ratios. Answer the questions about the following chemical reaction. sodium + iron(III) oxide →Y sodium oxide + iron 6Na(s) + →+ + 2Fe(s) 15. What is a mole ratio? 16.

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STUDV GUIDE In your textbook, read about why reactions stop and how to determine the limiting reactant. Study the diagram showing a chemical reaction and the chemical equation that repre- sents the reaction. Then complete the table. Show your calculations for questions 25—27 ... Study Guide 3. 12. 24. 27. 76 30 NO

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