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general.

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Justify your answer. Yes. Addition of an excess of 0.20 M  $\text{AgNO}_3$  (aq) will precipitate all of the I ion present in the solution because  $\text{AgI}$  is insoluble, as evidenced by its low value of  $K_{sp}$ . 1 point is earned for the correct answer

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with a valid justification. (ii) The student only has access to one KI

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Time—90 minutes CALCULATORS ARE

NOT ALLOWED FOR SECTION 1. Note:

For all questions, assume that the

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temperature is 298 K, the pressure is 1.0 atm, and solutions are aqueous unless otherwise specified. Directions: Each of the questions or incomplete statements below is followed by four suggested answers or

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01:00 pm. (The duration of paper was 3 hours) NEET exam 2014 paper has been published in Hindi, English. The exam paper is of objective type. There are total 180 questions. The total duration is of 3 hours for this paper. Physics Notes  
Chemistry Notes Biology Notes

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## Answers - AglaSem Admission

DRAFT 5/7/2014 10:27:00 PM 4 Question

4 (a)  $P V = n R T$   $(1.04)(1.00) = n$

$(0.08206)(1100)$   $n = 0.0115$  moles (1)

(b) Disagree. The reaction is at equilibrium so there must be some  $\text{CaCO}_3(\text{s})$  left in the container. (It is entirely possible to answer this via stoichiometry, proving that if there was

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total decomposition, that the moles and

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