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the SI unit representing 6.02×10^{23} representative particles of a substance. the temperature and pressure at which one mole of gas occupies a volume of 22.4 L. equal volumes of gases at the same temperature and pressure contain equal numbers of particles.

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SECTION 10.1 THE MOLE: A

MEASUREMENT OF MATTER (pages 287-296) This section defines the mole and explains how the mole is used to measure matter. It also teaches you how to calculate the mass of a mole of any

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substance. Measuring Matter (pages 287-289) 1.

SECTION 10.1 THE MOLE: A MEASUREMENT OF MATTER (pages 287-296)

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6.02 • 1023 representative particles of a substance molecule formula unit atom count the matter; measure the mass or weight; measure the volume They are questions about the amount of a substance and are similar to questions scientists ask.

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Chapter 10 Chemical Quantities 241.
Section Review. Objectives. •Relate Avogadro's number to a mole of a substance. •Calculate the mass of a mole of any substance. •Describe methods of measuring the amount of something. •Compare and contrast the atomic mass of an element and its molar mass. Vocabulary Key Equations.

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Objectives Vocabulary Key Equations

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