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#### 3 PROBLEMS Write the

answer on the line to

the left. Show all your

work in the space

provided. 1. 88% The

actual yield of a

reaction is 22 g and

the theoretical yield is

25 g. Calculate the

percentage yield. 2.

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6.0 mol of  $N_2$  are  
mixed with 12.0 mol of  
 $H_2$  according to the  
following equation:  $N_2(g) + 3H_2(g) \rightarrow 2NH_3(g)$

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Modern Chemistry 77  
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answer on the line to  
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provided. 1.          The actual yield of a reaction is 22 g and the theoretical yield is 25 g. Calculate the percentage yield. 2. 6.0 mol of  $N_2$  are mixed with 12.0 mol of H

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Stoichiometry.

SECTION 1. SHORT

ANSWER Answer the following questions in the space provided. 1.

\_\_\_\_\_ The coefficients in a chemical equation represent the (a) masses in grams of all

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## Stoichiometry

reactants and products. (b) relative number of moles of reactants and products. (c) number of atoms of each element in each compound in a reaction.

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3 PROBLEMS Write the  
answer on the line to  
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provided. 1. 88% The  
actual yield of a

reaction is 22 g and  
the theoretical yield is  
25 g. Calculate the  
percentage yield. 2.

6.0 mol of  $N_2$  are  
mixed with 12.0 mol of  
 $H_2$  according to the  
following

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3 PROBLEMS Write the answer on the line to the left Show all your work in the space provided

1 88% The actual yield of a reaction is 22 g and the theoretical yield is 25 g Calculate the percentage yield

2 60 mol of  $N_2$  are mixed with 120 mol of H

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CHAPTER 9 REVIEW.

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## Stoichiometry

### SECTION 9.2. Answer

PROBLEMS Write the answer on the line to the left. Show all your work in the space provided.

1. The following equation represents a laboratory preparation for oxygen gas:  $2\text{KClO}_3(\text{s}) \rightarrow 2\text{KCl}(\text{s}) + 3\text{O}_2(\text{g})$  How many grams of  $\text{O}_2$  form if 3.0 mol of  $\text{KClO}_3$  are totally consumed?

2. Given the following equation

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## Stoichiometry

### Section 3 Answer

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the line to the left show

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space provided 1 88  
the actual yield of a  
reaction is 22 g

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SECTION 2. PROBLEMS

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Write the answer on  
the line to the left.

Show all your work in  
the space provided. 1.

The following equation  
represents a laboratory  
preparation for oxygen  
gas:  $2\text{KClO}_3(\text{s}) \rightarrow 2\text{KCl}(\text{s}) + 3\text{O}_2(\text{g})$

CHAPTER 9 REVIEW ...

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Reaction stoichiometry  
uses molar  
relationships to  
determine the amounts

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of unknown reactants  
or products from the  
amounts of known  
reactants or products.

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100. of lithium carbonate  
Li<sub>2</sub>CO<sub>3</sub> — h. The oxygen  
gas produced in part  
has density of 1.43  
g/L calculate the volume  
of this gas. 76

## STOICHIOMETRY

## MODERN CHEMISTRY a.

— 81 g 6. A car air bag  
requires 70. L of  
nitrogen gas ...

**Date. FCHAPJ**

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composition  
stoichiometry. deals  
with the mass  
relationships of  
elements in  
compounds. ...

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rights In your

notebook, solve the

following problems.

SECTION 9.1 THE

ARITHMETIC OF

EQUATIONS Use the

3-step problem-solving

approach you learned

in Chapter 4. 1. An

apple pie needs 10

large apples, 2 crusts

(top and bottom), and

1 tablespoon of

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Stoichiometry  
cinnamon.

## Section 3 Answer

### **9 Stoichiometry Practice Problems**

CHAPTER 9 REVIEW  
Stoichiometry SECTION

3 PROBLEMS Write the  
answer on the line to  
the left Show all your  
work in the space

provided 1 88% The  
actual yield of a  
reaction is 22 g and  
the theoretical yield is  
25 g Calculate the  
percentage yield 2 60  
mol of N<sub>2</sub> are mixed

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Stoichiometry  
Section 2 Answer  
Key

with 120 mol of  $H_2$   
according to the  
following equation:  $N_2(g) + 3H_2(g)$

## **Download Chapter 9 Review Stoichiometry Section 2 Work**

Chapter 9 focuses on reaction stoichiometry: using a balanced chemical equation to calculate the number of grams, moles, or particles of reactants/products

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involved in a chemical reaction. Students had an introduction to composition stoichiometry in Chapter 3 and will now move on to some more difficult problems.

## **Stoichiometry Worksheet Answers Chapter 9**

CHAPTER 9 REVIEW  
Stoichiometry SECTION  
2 PROBLEMS Write the  
answer on the line to  
the left. Show all your

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work in the space  
provided.

1. 4.5 mol

The following equation  
represents a laboratory

preparation for oxygen  
gas:  $2\text{KClO}_3(\text{s}) \rightarrow$

$2\text{KCl}(\text{s}) + 3\text{O}_2(\text{g})$  How

many moles of  $\text{O}_2$

form if 3.0 mol of  $\text{KClO}_3$

are totally

consumed? ...

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