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in Atoms. Section 5.1 -
Models of the Atom.

The Rutherford's model
of the atom did not
explain how an atom
can emit light or the
chemical properties of
an atom. Plum Pudding
Model Rutherford's
Model. The Bohr Model.

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representing the
atomic nucleus and
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models of light. You
will describe how the
frequency of light
emitted by an atom is
a unique characteristic
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compare and contrast

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all electromagnetic

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waves in a vacuum is the same, waves can have different wavelengths and frequencies. As you can see from the equation on the previous page, wavelength and frequency are inversely related; in other words, as one quantity increases, the other decreases.

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CHE Chapter 5:

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consists of an element's symbol, representing the atomic nucleus and inner-level electrons, that is surrounded by dots, representing the atom's valence

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MODELS OF THE ATOM
(pages 127-132) This
section summarizes
the development of
atomic theory. It also
explains the
significance of
quantized energies of

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electrons as they relate to the quantum mechanical model of the atom. The Development of Atomic Models (pages 127-128) 1.

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Electrons occupy equal energy orbitals singly before beginning to

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pair up, because according to Hund's rule the most stable arrangement of electrons occurs when electrons with the same spin is as large as possible.

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02:25. Problem 1

Objects get their colors from reflecting only certain wavelengths when hit with white light. Light reflected from a green leaf is found to have a wavelength of $4.90 \times 10^{-7} \text{ m}$

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