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way. Mcqs Solution And Colligative Properties Mcq Which observation(s) reflect(s) colligative properties? (I) A 0.5 m NaBr solution has a higher vapor pressure than a 0.5 m BaCl<sub>2</sub> solution. (II) A 0.5 m NaOH solution freezes at a lower temperature than pure water. (III) Pure water freezes at a higher temperature than pure methanol.

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Colligative properties depend on . (i) the nature of the solute particles dissolved in solution. (ii) the number of solute particles in solution. (iii) the physical properties of the solute particles dissolved in solution. (iv) the nature of solvent particles. Which of the following aqueous solutions should have the highest boiling point? (i) 1 ...

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In chemistry, a solution is a special type of homogeneous mixture which is composed of 2 or more substances. The term aqueous solution is used for the solutions where one of the

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solvents is water. In such a mixture, a solute is a substance dissolved in another substance, known as a solvent. Chemistry Solutions MCQs. 1.

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## Mcqs Solution And Colligative Properties Mcq

Questions MCQ Test Solutions & Colligative Properties MCQs. View source. History Talk (0) Share. Contents Level 1 . Level 2 . Level 3 . Q.1) The lowering of vapour pressure of a solution is 15 mm Hg if mole fraction of solute is 0.2 . What is the mole fraction of solute if the lowering of vapour pressure is

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Which observation(s) reflect(s) colligative properties? (I) A 0.5 m NaBr solution has a higher vapor pressure than a 0.5 m BaCl<sub>2</sub> solution. (II) A 0.5 m NaOH solution freezes at a lower temperature than pure water. (III) Pure water freezes at a higher temperature than pure methanol. (a) only I (b) only II (c) only III (d) I and II (e) I and III 7.

## Solutions - Texas A&M University

Aqueous solutions are found throughout our world, and their chemistry depends in part on how much of a dissolved substance is in them. This module explores how substances dissolve, why some substances don't dissolve, and how we express the concentration of a solution. The module describes the forces that hold particles together and interactions that keep dissolved particles apart.

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